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English advance in Asia. 1600: Michael the Brave unites the three principalities: Wallachia, Moldavia and Transylvania after the Battle of Selimbar in 1599. For later events, see Timeline of the 17th century. Polybius The Histories translated into Italian, English, German and French.[20] Mississippian culture disappears. Medallion rug, variant Star Ushak style, Anatolia (modern Turkey), is made. It is now kept at the Saint Louis Art Museum. Hernan Cortes (1495–1547) Henry VIII, (1491–1547) King of England and Ireland Don Fernando Alvarz de Toledo (1507–1582) Suleiman the Magnificent, Sultan of the Ottoman Empire (1520–1566) Ivan IV the Terrible (1530–1584) Oda Nobunaga (1534–1582) Sir Francis Drake (c. 1540 – 1596) Alberico Gentili, (1552–1608) the Father of international law Philip II of Spain, King of Spain (1556–1598) Akbar the Great, Mughal emperor (1556–1605) Related article: List of 16th century inventions. The Columbian Exchange introduces many plants, animals and diseases to the Old and New Worlds. Introduction of the spinning wheel revolutionizes textile production in Europe. The letter J is introduced into the English alphabet. 1500: First portable watch is created by Peter Henlein of Germany.The Iberian Union in 1598, under Philip II, King of Spain and Portugal 1513: Juan Ponce de León sights Florida and Vasco Núñez de Balboa sights the eastern edge of the Pacific Ocean. 1519–1522: Ferdinand Magellan and Juan Sebastián Elcano lead the first circumnavigation of the world. 1519–1540: In America, Hernando de Soto expeditions map the Gulf of Mexico coastline and bays. 1525: Modern square root symbol (√) 1540: Francisco Vázquez de Coronado sights the Grand Canyon. 1541–42: Francisco de Orellana sails the length of the Amazon River. 1542–43: Firearms are introduced into Japan by the Portuguese. 1543: Copernicus publishes his theory that the Earth and the other planets revolve around the Sun 1545: Theory of complex numbers is first developed by Gerolamo Cardano of Italy. 1558: Camera obscura is first used in Europe by Giambattista della Porta of Italy. 1559–1562: Spanish settlements in Alabama/Florida and Georgia confirm dangers of hurricanes and local native warring tribes. 1565: Spanish settlers outside New Spain (Mexico) colonize Florida's coastline at St. Augustine. 1565: Invention of the graphite pencil (in a wooden holder) by Conrad Gesner. Modernized in 1812. 1568: Gerardus Mercator creates the first Mercator projection map. 1572: Supernova SN 1572 is observed by Tycho Brahe in the Milky Way. 1582: Gregorian calendar is introduced in Europe by Pope Gregory XIII and adopted by Catholic countries. c. 1583: Galileo Galilei of Pisa, Italy identifies the constant swing of a pendulum, leading to development of reliable timekeepers. 1585: earliest known reference to the 'sailing carriage' in China. 1589: William Lee invents the stocking frame. 1591: First flush toilet is introduced by Sir John Harrington of England, the design published under the title 'The Metamorphosis of Ajax'. 1593: Galileo Galilei invents a thermometer. 1596: William Barents discovers Spitsbergen. 1597: Opera in Florence by Jacopo Peri. Entertainment in the 16th century ^ a b Modern reference works on the period tend to follow the introduction of the Gregorian calendar for the sake of clarity; thus NASA's lunar eclipse catalogue states "The Gregorian calendar is used for all dates from 1582 Oct 15 onwards. Before that date, the Julian calendar is used." For dates after 15 October 1582, care must be taken to avoid confusion of the two styles. ^ de Vries, Jan (14 September 2009). "The limits of globalization in the early modern world". *The Economic History Review*. 63 (3): 710–733. CiteSeerX 10.1.1.186.2862. doi:10.1111/j.1468-0289.2009.00497.x. JSTOR 40929823. S2CID 219969360. SSRN 1635517. ^ Singh, Sarina; Lindsay Brown; Paul Clammer; Rodney Cocks; John Mock (2008). *Pakistan & the Karakoram Highway*. Vol. 7, illustrated. Lonely Planet. p. 137. ISBN 978-1-74104-542-0. Retrieved 23 August 2010. ^ Babur (2006). *Babur Nama*. Penguin Books. p. vii. 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An Encyclopedia of World History (5th ed. 1973); highly detailed outline of events online free Media related to 16th century at Wikimedia Commons Timelines of 16th century events, science, culture and persons Retrieved from " 4 The following pages link to 16th century External tools (link count transclusion count sorted list) · See help page for transcluding these entries Showing 50 items. View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500)Bagpipes (links | edit) List of decades, centuries, and millennia (links | edit) Fashion (links | edit) Giovanni Boccaccio (links | edit) History of Mali (links | edit) History of Mauritius (links | edit) Post office (links | edit) Snare drum (links | edit) Republican Party (United States) (links | edit) 20th century (links | edit) 15th century (links | edit) 17th century (links | edit) 18th century (links | edit) 1624 (links | edit) 1626 (links | edit) 1642 (links | edit) 1661 (links | edit) 1608 (links | edit) 1492 (links | edit) 14th century (links | edit) 1st century (links | edit) 13th century (links | edit) 4th century (links | edit) 12th century (links | edit) 11th century (links | edit) 1564 (links | edit) 1648 (links | edit) 1572 (links | edit) 1623 (links | edit) 1662 (links | edit) 1490s (links | edit) 1640s (links | edit) 1597 (links | edit) 1690 (links | edit) 1688 (links | edit) 7th century (links | edit) 10th century (links | edit) 9th century (links | edit) 8th century (links | edit) 5th century (links | edit) 3rd century (links | edit) 2nd century (links | edit) 1573 (links | edit) 1570s (links | edit) 1574 (links | edit) 1436 (links | edit) 1476 (links | edit) 1542 (links | edit) 1540s (links | edit) View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500) Retrieved from " WhatLinksHere/16th century" The modern periodic table is based on the law that an element's properties are a periodic function of its atomic number. These properties are related to the elements' electronic configuration. As we move across a period from left to right or down the group, we notice a common trend in properties. This property trend is referred to as periodic properties. Atomic size, metallic character, non-metallic character, ionisation potential, electron affinity, and electronegativity are all important periodic properties. What is the Periodic Law? Periodic trends are founded on periodic law. The chemical elements are enumerated in order of increasing atomic number, and main properties undergo cyclic changes, according to the periodic law. At regular intervals, elements with comparable chemical characteristics reappear. Dmitri Mendeleev established this principle. He also asserted that the periodic chart was based on different physical and chemical properties of elements, not merely atomic weights. The recurrence of features was later discovered to be attributable to the recurrence of comparable electronic configurations in the outer shells of atoms. Periodic Table Trends The periodic trends are based on the Periodic Law, which states that if the chemical elements are listed in increasing atomic number order, many of their properties undergo cyclical changes, with elements with similar properties recurring at regular intervals. Many of the physical and chemical properties of lithium, such as its vigorous reactivity with water, recur in sodium, potassium, and caesium after arranging elements in increasing atomic numbers. Mendeleev discovered this principle in 1871, following a number of investigations throughout the nineteenth century. Mendeleev also proposed an elemental periodic system based not only on atomic weights but also on the chemical and physical properties of the elements and their compounds. Henry Moseley discovered in 1913 that periodicity is determined by the atomic number rather than the atomic weight. Lothar Meyer presented his table following Mendeleev, but he disagreed with Mendeleev's Periodic law. Initially, there was no theoretical explanation for the Periodic Law, and it was only used as an empirical principle; however, with the development of quantum mechanics, it became possible to understand the theoretical basis for the Periodic Law. When elements are listed in increasing atomic number order, the periodic recurrence of elements with similar physical and chemical properties results directly from the periodic recurrence of similar electronic configurations in respective atoms' outer shells. One of the most significant events in the history of chemical science was the discovery of the Periodic Law. Almost every chemist employs and continues to employ the Periodic Law. The Periodic Law also resulted in the creation of the periodic table, which is now widely used in a variety of fields. Atomic size The atomic radius is the distance between the nucleus's centre and the atom's outermost shell. As we progress from one period to the next, the atomic size of a group grows due to the addition of shells. Over time, the atomic size decreases while the number of shells remains constant and the nuclear charge increases. This causes electrons to be drawn from the outermost shell towards the nucleus, reducing its size. Metallic character Metals are the elements that lose electrons to form cations. The metallic character increases as we move down the group because atomic size increases, resulting in easy electron loss. It, on the other hand, decreases over time as we move from left to right. This occurs because the nuclear charge increases, making it more difficult for an atom to lose electrons. Non-metallic character Non-metals are elements that have a proclivity to gain electrons. Moving across a period increases the tendency to gain electrons due to an increase in nuclear charge and a decrease in atomic size. As a result, the nonmetallic character increases over time. The non-metallic character decreases as we move down the group due to the increase in atomic size. Ionization potential The ionisation potential is defined as the amount of energy required to remove an electron from the outermost shell of a gaseous atom and convert it into a positively charged gaseous ion. The periodic properties of ionisation potential increase as the atomic size decreases over time due to an increase in nuclear charge. The ionisation potential decreases as we move down the group due to the increase in atomic size. Factors That Affect the Ionization Energy Levels Nuclear Charge: The force of attraction between the nucleus and valence electrons decreases as the nuclear charge decreases, resulting in decreased ionisation energy.Shielding Effect: Shielding effect increases as nuclear charge grows, hence as the shielding effect increases, so does ionisation energy.Atomic Radius: The force of attraction between the nucleus and valence electrons reduces as the atomic radius grows. As a result, as the atomic radius increases, the ionisation reduces.Half-Filled Valence Shells: Ionization energy is high in pseudo-filled or half-filled valence shells. The periodic trend does not apply to any of the elements in the oxygen and boron families. They take a little less energy than the standard trend. Electron affinity A certain amount of energy is released when electrons are added to a neutral gaseous atom. This is referred to as electron affinity. Electron affinity causes the formation of a negative ion or anion. On the periodic table, it increases from left to right and decreases from top to bottom. Electron affinity decreases as atomic size increases, and vice versa. The electron affinity is also affected by the screening effect and the reactivity of non-metals. Melting Point The melting point of an element is the amount of energy required to change the state of an element from solid to liquid. Which essentially means severing a few ties. As a result, the melting point rises as the strength of the bond between the atoms increases. Boiling Point The boiling point is determined by the amount of heat required to convert a liquid to a gaseous state, just as the melting point is determined by the strength of the bonds between atoms. Electronegativity The ability of an atom to attract a pair of electrons is referred to as electronegativity. The electronegativity of elements is affected by the size of an atom and its nuclear charge. It increases as we move from left to right across the periodic table and decreases as we move from top to bottom. Fluorine has a higher electronegative value, while caesium has a lower electronegative value. It also differs between metals and non-metals. Nonmetals have a higher electronegative potential than metals. It also aids in the identification of the types of bonds formed between the elements. The group 13 elements are an exception, and electronegativity rises from aluminium to thallium as a result. In addition, tin has a stronger electronegativity than lead in group 14. Metallic character Metallic character is defined as the characteristics associated with the metals found on the periodic table. Metallic lustre, hardness, malleability, thermal conductivity, and other properties are examples of these characteristics. The elements on the left side of the periodic table have a more metallic character. It decreases from left to right due to electron addition and increases from top to bottom due to electron removal. When the atomic number increases, so do the ability to lose electrons. Non-metallic character The periodic table shows that the nonmetallic character of elements increases from left to right and decreases from top to bottom. The elements with this property do not have metallic properties. Shielding Effect It can be characterized as the inner electrons resisting an outer electron. It can also be used to describe how many nuclei are capable of controlling electrons in the outer shell. Because of the increasing shielding effect, the effective nuclear charge drops along with the group. The effective nuclear charge rises over time as the nuclear charge rises. Sample Questions Question 1: What is the modern periodic table based on? Answer: The modern periodic table is based on the law that an element's properties are a periodic function of its atomic number. These properties are related to the elements' electronic configuration. Question 2: What is the effect on metallic character as we move in a periodic table? Answer: The metallic character increases as we move down the group because atomic size increases, resulting in easy electron loss. It, on the other hand, decreases over time as we move from left to right. Question 3: Define electron affinity. Answer: A certain amount of energy is released when electrons are added to a neutral gaseous atom. This is referred to as electron affinity. Question 4: What is the effect on electronegativity as we move in a periodic table? Answer: The electronegativity of elements is affected by the size of an atom and its nuclear charge. It increases as we move from left to right across the periodic table and decreases as we move from top to bottom. Question 5: What are some metallic characteristics? Answer: Metallic character is defined as the characteristics associated with the metals found on the periodic table. Metallic lustre, hardness, malleability, thermal conductivity, and other properties are examples of these characteristics.