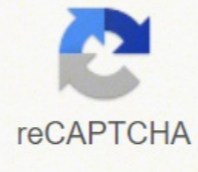
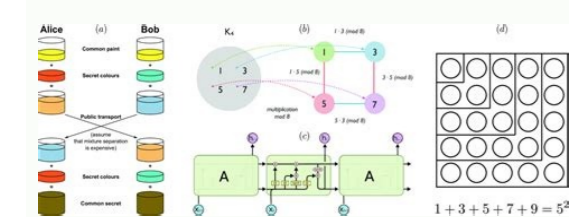




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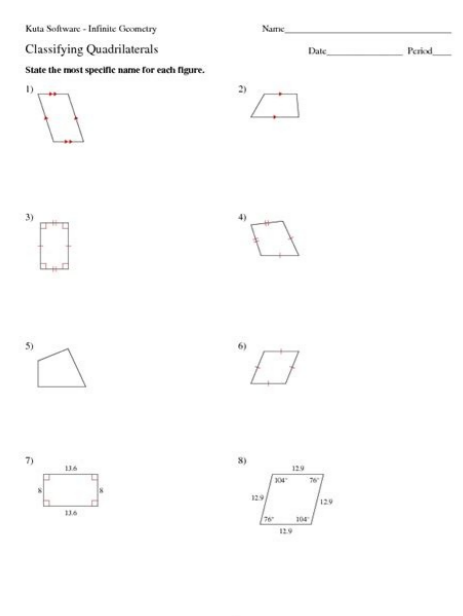
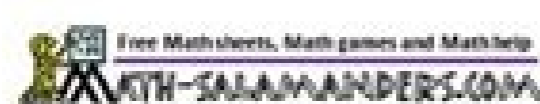


EXPLORING 2D SHAPES SHEET 5 UK ANSWERS

ALWAYS, SOMETIMES OR NEVER TRUE?

STATEMENT	ALWAYS, SOMETIMES OR NEVER TRUE?	REASONING
1) A trapezium has exactly one right angle.	ALWAYS SOMETIMES NEVER	A trapezium either has none, two or four right angles. (*)
2) The diameter of a circle is half the radius.	ALWAYS SOMETIMES NEVER	The radius of a circle is always equal to half the diameter, not the other way round!
3) The opposite angles of a parallelogram are equal.	ALWAYS SOMETIMES NEVER	This is always true. Equal angles: The red angles are equal. The blue angles are equal.
4) A kite has a line of symmetry.	ALWAYS SOMETIMES NEVER	All kites have at least one line of symmetry.
5) A triangle has exactly two mirror lines.	ALWAYS SOMETIMES NEVER	A triangle can have none, one or three lines of symmetry but never two!
6) A parallelogram has exactly two lines of symmetry.	ALWAYS SOMETIMES NEVER	Parallelograms can have 0, 2 or 4 lines of symmetry. This rectangle, which is a parallelogram, has two.

(*) Most mathematicians define a trapezium as have at least one pair of parallel lines (therefore including squares, rectangles, parallelograms, etc. as special cases of a trapezium).



Acids and Bases Worksheet 1

1. Finish the following reactions when acids are added to water.

- $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow$
- $\text{HCl} + \text{OH}^- \rightarrow$
- $\text{NH}_3 + \text{H}_2\text{O} \rightarrow$
- $\text{HCl} + \text{NH}_3 \rightarrow$

2. What are the conjugate bases of these acids?

original acid	conjugate base
HNO_3	
H_2O	
H_3O^+	
H_2SO_4	
HBr	
HCO_3^-	

3. What are the conjugate acids of these bases?

original base	conjugate acid
OH^-	
H_2O	
HCO_3^-	
SO_4^{2-}	
ClO_4^-	

4. Which of the following represent conjugate acid-base pairs?

- $\text{H}_2\text{O}, \text{H}_3\text{O}^+$
- $\text{OH}^-, \text{HNO}_3$
- $\text{H}_2\text{SO}_4, \text{SO}_4^{2-}$
- $\text{HC}_2\text{H}_3\text{O}_2, \text{C}_2\text{H}_3\text{O}_2^-$

5. Calculate the $[\text{H}^+]$ in a solution in which $[\text{OH}^-] = 2.0 \times 10^{-2} \text{ M}$. Is this solution acidic, neutral, or basic?

6. Calculate the $[\text{OH}^-]$ in a solution in which $[\text{H}^+] = 3.99 \times 10^{-9} \text{ M}$. Is this solution acidic, neutral, or basic?

7. Convert each of the following into pH. Identify each as an acidic, neutral, or basic.

- 0.0015 M H^+

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